CHEM 3.7 Practise Electrochemistry Questions

a) Using the following reduction potentials discuss which species would undergo spontaneous reaction with **bromine, Br2**, under standard conditions. Justify your answer and write balanced equations for the reactions occurring.

*E*°(Br2/Br–) = +1.07 V, *E*°(Sn4+/Sn2+) = +0.15 V,

*E*°(Cu2+/Cu) = +0.34 V, *E*°(H2O2/H2O) = +1.77 V

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b) Cobalt is a transition metal that exists in both the +2 and +3 oxidation states.

 A piece of cobalt metal is reacted with acidified **potassium dichromate** solution.

 Using the relevant reduction potentials, determine if the cobalt ion produced in this reaction is Co2+ or Co3+.

*E*°(Co3+/Co2+) = +1.82 V *E*°(Co2+/Co) = –0.28 V *E*°(Cr2O72-/Cr3+) = +1.36 V

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c) The basis for all batteries is electrochemical cells. Below is a partly completed drawing of the cell: Fe⏐Fe2+⏐⏐Sn2+⏐Sn

*E*°(Sn2+/Sn) = –0.14 V and *E*°(Fe2+/Fe) = –0.44 V



(i)On the diagram above

complete the cell including the external circuit,

fully label each half-cell (including charge of each electrode),

identify where oxidation and reduction occur

give balanced equations for the half-reactions occurring at each electrode

indicate the direction of flow of charged particles in both the internal and external circuits

(ii) Discuss the features that you have drawn on the cell-diagram above, using the electrode potentials to support your answer.

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